



# Periodic Table PRO

The chemistry lab in your pocket.





Our mission is to educate younger generations through technology. Chemistry is one of the fascinating matters and its value is given by the quality of the educational act.

We strongly believe that education must include elements that stimulate visual cognitive processes and also inspire students to enjoy the process of learning.

Periodic Table PRO provides concise information for all chemical elements, a virtual laboratory for conducting fun experiments, as well as an environment dedicated to testing the knowledge gained in chemistry classes.



# Periodic Table of Elements



Groups 2   **4**   6

## Group 01

7 elements

1 <sup>1</sup>  
**H**  
Hydrogen  
1.0079

3 <sup>2</sup><sub>1</sub>  
**Li**  
Lithium  
6.941

11 <sup>2</sup><sub>8</sub><sub>1</sub>  
**Na**  
Sodium  
22.990



+4

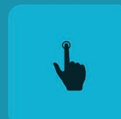
## Group 02

6 elements

4 <sup>2</sup><sub>2</sub>  
**Be**  
Beryllium  
9.0122

12 <sup>2</sup><sub>8</sub><sub>2</sub>  
**Mg**  
Magnesium  
24.305

20 <sup>2</sup><sub>8</sub><sub>8</sub><sub>2</sub>  
**Ca**  
Calcium  
40.078



+3

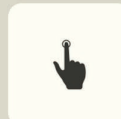
## Group 03

4 elements

21 <sup>2</sup><sub>8</sub><sub>9</sub><sub>2</sub>  
**Sc**  
Scandium  
44.956

39 <sup>2</sup><sub>8</sub><sub>18</sub><sub>9</sub><sub>2</sub>  
**Y**  
Yttrium  
88.906

57 <sup>2</sup><sub>8</sub><sub>18</sub><sub>18</sub><sub>9</sub><sub>2</sub>  
**La**  
Lanthanum  
138.91



+1

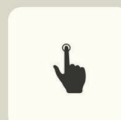
## Group 04

6 elements

22 <sup>2</sup><sub>8</sub><sub>10</sub><sub>2</sub>  
**Ti**  
Titanium  
47.867

40 <sup>2</sup><sub>8</sub><sub>18</sub><sub>10</sub><sub>2</sub>  
**Zr**  
Zirconium  
91.224

72 <sup>2</sup><sub>8</sub><sub>18</sub><sub>32</sub><sub>10</sub><sub>2</sub>  
**Hf**  
Hafnium  
178.49



+3

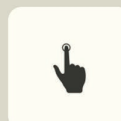
## Group 05

6 elements

23 <sup>2</sup><sub>8</sub><sub>11</sub><sub>2</sub>  
**V**  
Vanadium  
50.942

41 <sup>2</sup><sub>8</sub><sub>18</sub><sub>12</sub><sub>1</sub>  
**Nb**  
Niobium  
92.906

73 <sup>2</sup><sub>8</sub><sub>18</sub><sub>32</sub><sub>11</sub><sub>2</sub>  
**Ta**  
Tantalum  
180.95



+3

EN



# Periodic Table of Elements



List

2 4 6

|  |                                       |  |  |
|--|---------------------------------------|--|--|
| 1<br><b>H</b><br>Hydrogen<br>1.0079    | 2<br><b>He</b><br>Helium<br>4.0026    | 3<br><b>Li</b><br>Lithium<br>6.941     | 4<br><b>Be</b><br>Beryllium<br>9.0122  |
| 5<br><b>B</b><br>Boron<br>10.811       | 6<br><b>C</b><br>Carbon<br>12.011     | 7<br><b>N</b><br>Nitrogen<br>14.007    | 8<br><b>O</b><br>Oxygen<br>15.999      |
| 9<br><b>F</b><br>Fluorine<br>18.998    | 10<br><b>Ne</b><br>Neon<br>20.180     | 11<br><b>Na</b><br>Sodium<br>22.990    | 12<br><b>Mg</b><br>Magnesium<br>24.305 |
| 13<br><b>Al</b><br>Aluminium<br>26.982 | 14<br><b>Si</b><br>Silicon<br>28.086  | 15<br><b>P</b><br>Phosphorus<br>30.974 | 16<br><b>S</b><br>Sulfur<br>32.065     |
| 17<br><b>Cl</b><br>Chlorine<br>35.453  | 18<br><b>Ar</b><br>Argon<br>39.948    | 19<br><b>K</b><br>Potassium<br>39.098  | 20<br><b>Ca</b><br>Calcium<br>40.078   |
| 21<br><b>Sc</b><br>Scandium<br>44.956  | 22<br><b>Ti</b><br>Titanium<br>47.867 | 23<br><b>V</b><br>Vanadium<br>50.942   | 24<br><b>Cr</b><br>Chromium<br>51.996  |
| 25<br><b>Mn</b><br>Manganese<br>54.938 | 26<br><b>Fe</b><br>Iron<br>55.845     | 27<br><b>Co</b><br>Cobalt<br>58.933    | 28<br><b>Ni</b><br>Nickel<br>58.693    |

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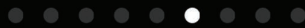
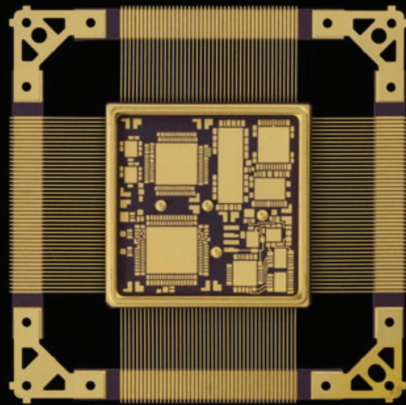


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# Gold



Gold is one of the few elements you can find just lying on the ground. This one-ounce nugget of pure gold was found in Alaska in 1890 by Hogamorth Marion.



👉 Swipe for more details

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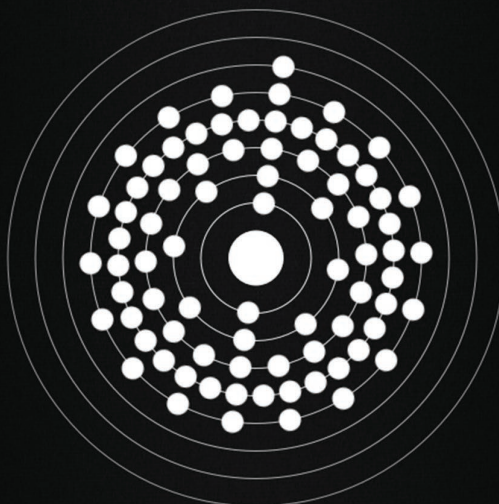
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# Gold



## 01. Overview

|               |                        |
|---------------|------------------------|
| Symbol        | Au                     |
| Atomic number | 79                     |
| Atomic weight | 197                    |
| Density       | 19.3 g/cm <sup>3</sup> |
| Melting point | 1064.18 °C             |
| Boiling point | 2856 °C                |
| Electrons     | 79                     |
| Protons       | 79                     |
| Neutrons      | 118                    |



## 02. Thermal properties

|                        |            |
|------------------------|------------|
| Phase                  | Solid      |
| Melting point          | 1064.18 °C |
| Boiling point          | 2856 °C    |
| Absolute melting point | 1337.33 K  |
| Absolute boiling point | 3129 K     |

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Element details - PDF




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Experiments



 You have found 0 of 11 secrets.

Chernobyl



The "Barking dog"



Cesium and water



Nitrogen triiodide



The Cannon



Rubber egg



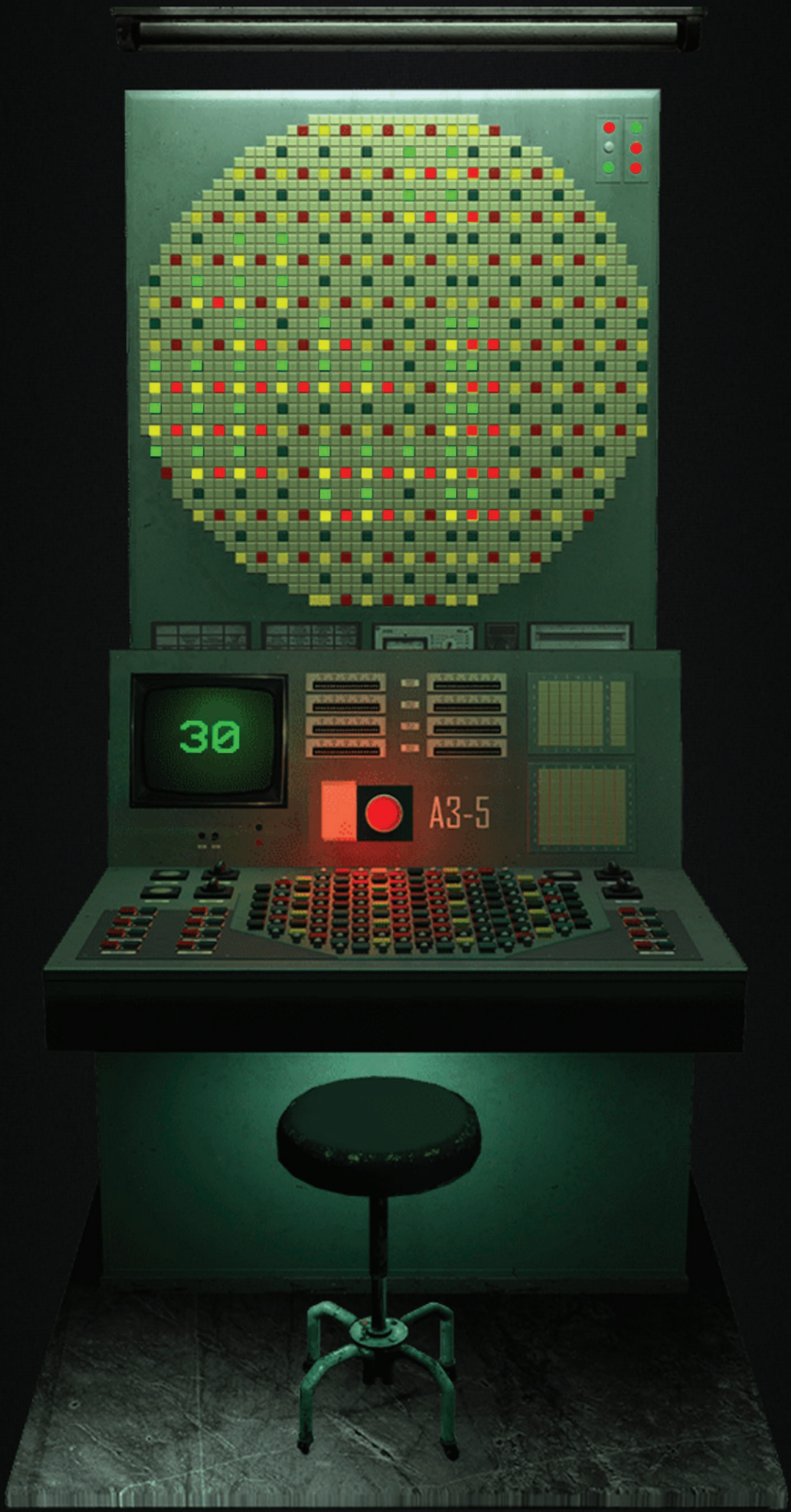
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Chernobyl



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Soft drink + lollies

Soft drink



Lollies



Soft drink + lollies

? Help

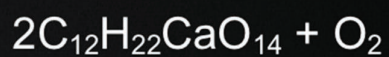
In this experiment you can view the reaction between a soft drink and a bunch of lollies. To make the experiment, drag the elements that are pulsating into the glass at the center of the screen.

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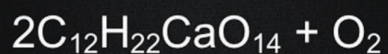
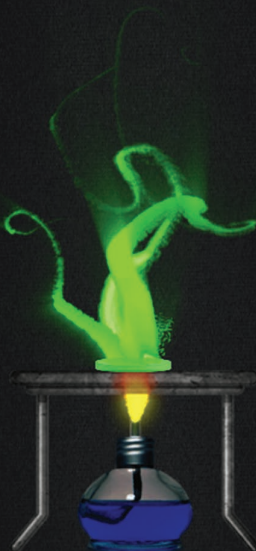
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Calcium gluconate



Fire



? Help

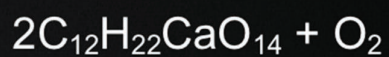
In this experiment you can view the reaction between calcium gluconate and fire. To make the experiment, drag the elements that are pulsating into the glass at the center of the screen.

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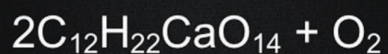
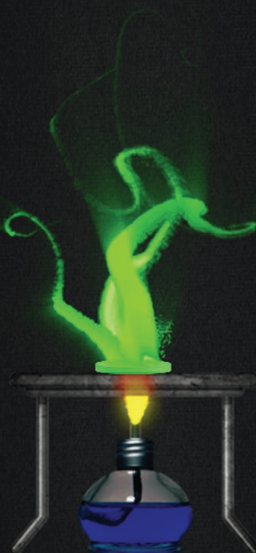
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Calcium gluconate



Fire



? Help

In this experiment you can view the reaction between calcium gluconate and fire. To make the experiment, drag the elements that are pulsating into the glass at the center of the screen.

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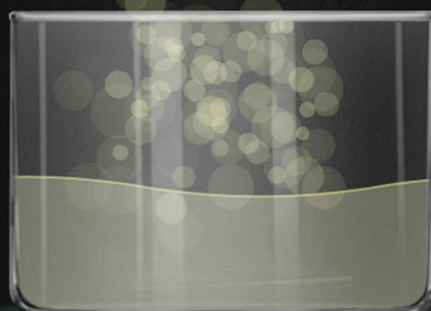
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HCl + Mg

Hydrochloric acid



Magnesium



HCl + Mg

? Help

In this experiment you can view the reaction between the hydrochloric acid and magnesium solution. To make the experiment, drag the elements that are pulsating into the glass at the center of the screen.

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# Periodic Table of Elements

|                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                                 |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                             |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                            |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                            |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                            |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
|--------------------------|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|---------------------------|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|---------------------------|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|---------------------------------|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|---------------------------|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|---------------------------|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|---------------------------|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|---------------------------|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|---------------------------|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|-----------------------------|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|----------------------------|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|----------------------------|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--------------------------|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--------------------------|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--------------------------|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|----------------------------|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|---------------------------|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--------------------------|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|
| 1                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 18                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                                 |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                             |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                            |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                            |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                            |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| 1                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 2                         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                                 |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                             |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                            |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                            |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                            |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| H<br>Hydrogen<br>1.0079  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | He<br>Helium<br>4.0026    |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                                 |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                             |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                            |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                            |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                            |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| 3                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 4                         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 5                         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 6                               |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 7                         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 8                         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 9                         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 10                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 11                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 12                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 13                         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 14                         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 15                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 16                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 17                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 18                         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| Li<br>Lithium<br>6.941   |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Be<br>Beryllium<br>9.0122 |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | B<br>Boron<br>10.811      |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | C<br>Carbon<br>12.011           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | N<br>Nitrogen<br>14.007   |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | O<br>Oxygen<br>15.999     |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | F<br>Fluorine<br>18.998   |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Ne<br>Neon<br>20.180      |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                             |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                            |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                            |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                            |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| 11                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 12                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 13                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 14                              |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 15                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 16                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 17                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 18                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                             |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                            |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                            |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                            |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| Na<br>Sodium<br>22.990   |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Mg<br>Magnesium<br>24.305 |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Al<br>Aluminium<br>26.982 |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Si<br>Silicon<br>28.086         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | P<br>Phosphorus<br>30.974 |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | S<br>Sulfur<br>32.055     |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Cl<br>Chlorine<br>35.453  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Ar<br>Argon<br>39.948     |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                             |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                            |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                            |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                            |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                           |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| 19                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 20                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 21                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 22                              |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 23                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 24                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 25                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 26                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 27                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 28                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 29                         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 30                         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 31                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 32                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 33                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 34                         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 35                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 36                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| K<br>Potassium<br>39.098 |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Ca<br>Calcium<br>40.078   |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Sc<br>Scandium<br>44.956  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Ti<br>Titanium<br>47.867        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | V<br>Vanadium<br>50.942   |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Cr<br>Chromium<br>51.996  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Mn<br>Manganese<br>54.938 |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Fe<br>Iron<br>55.845      |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Co<br>Cobalt<br>58.933    |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Ni<br>Nickel<br>58.693      |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Cu<br>Copper<br>63.546     |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Zn<br>Zinc<br>65.38        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Ga<br>Gallium<br>69.723  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Ge<br>Germanium<br>72.64 |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | As<br>Arsenic<br>74.922  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Se<br>Selenium<br>78.96    |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Br<br>Bromine<br>79.904   |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Kr<br>Krypton<br>83.798  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| 37                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 38                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 39                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 40                              |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 41                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 42                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 43                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 44                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 45                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 46                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 47                         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 48                         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 49                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 50                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 51                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 52                         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 53                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 54                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| Rb<br>Rubidium<br>85.468 |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Sr<br>Strontium<br>87.62  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Y<br>Yttrium<br>88.906    |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Zr<br>Zirconium<br>91.224       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Nb<br>Niobium<br>92.906   |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Mo<br>Molybdenum<br>95.96 |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Tc<br>Technetium<br>[98]  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Ru<br>Ruthenium<br>101.07 |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Rh<br>Rhodium<br>102.91   |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Pd<br>Palladium<br>106.42   |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Ag<br>Silver<br>107.87     |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Cd<br>Cadmium<br>112.41    |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | In<br>Indium<br>114.82   |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Sn<br>Tin<br>118.71      |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Sb<br>Antimony<br>121.76 |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Te<br>Tellurium<br>127.60  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | I<br>Iodine<br>126.90     |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Xe<br>Xenon<br>131.29    |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| 55                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 56                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 57-71                     |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 72                              |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 73                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 74                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 75                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 76                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 77                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 78                          |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 79                         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 80                         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 81                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 82                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 83                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 84                         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 85                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 86                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| Cs<br>Cesium<br>132.91   |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Ba<br>Barium<br>137.33    |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Lanthanides               |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Hf<br>Hafnium<br>178.49         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Ta<br>Tantalum<br>180.95  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | W<br>Wolfram<br>183.84    |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Re<br>Rhenium<br>186.21   |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Os<br>Osmium<br>190.23    |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Ir<br>Iridium<br>192.22   |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Pt<br>Platinum<br>195.08    |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Au<br>Gold<br>196.97       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Hg<br>Mercury<br>200.59    |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Tl<br>Thallium<br>204.38 |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Pb<br>Lead<br>207.2      |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Bi<br>Bismuth<br>208.98  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Po<br>Polonium<br>[209]    |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | At<br>Astatine<br>[210]   |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Rn<br>Radon<br>[222]     |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| 87                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 88                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 89-103                    |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 104                             |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 105                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 106                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 107                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 108                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 109                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 110                         |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 111                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 112                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 113                      |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 114                      |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 115                      |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 116                        |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 117                       |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | 118                      |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |
| Fr<br>Francium<br>[223]  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Ra<br>Radium<br>[226]     |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Actinides                 |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Rf<br>Rutherfordium<br>[263.11] |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Db<br>Dubnium<br>[268]    |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Sg<br>Seaborgium<br>[271] |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Bh<br>Bohrium<br>[270]    |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Hs<br>Hassium<br>[269]    |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Mt<br>Meitnerium<br>[278] |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Ds<br>Darmstadtium<br>[281] |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Rg<br>Roentgenium<br>[281] |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Cn<br>Copernicium<br>[285] |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Nh<br>Nihonium<br>[286]  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Fl<br>Flerovium<br>[289] |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Mc<br>Moscovium<br>[285] |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Lv<br>Livermorium<br>[293] |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Ts<br>Tennessine<br>[294] |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  | Og<br>Oganesson<br>[294] |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |  |

### Legend

- Alkali Metals
- Alkali Earth Metals
- Transition Metals
- Lanthanides
- Actinides
- Other Metals
- Metalloids
- Non-Metals
- Halogens
- Noble Gases

Lanthanides

Actinides

EN



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79

Au

196.97

2  
8  
18  
32  
18  
1



Gold

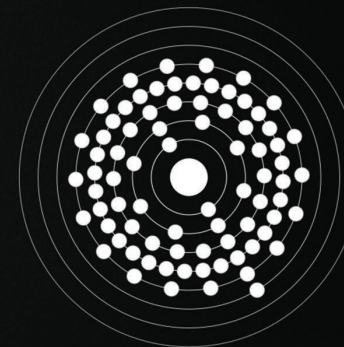
Atomic weight 197  
Density 19.3 g/cm<sup>3</sup>  
Melting point 1064.18 °C  
Boiling point 2856 °C

Gold is one of the few elements you can find just lying on the ground. This one-ounce nugget of pure gold was found in Alaska in 1890 by Hogamorth Marion.

Element details - PDF

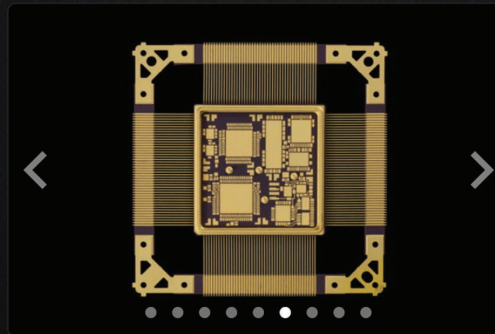
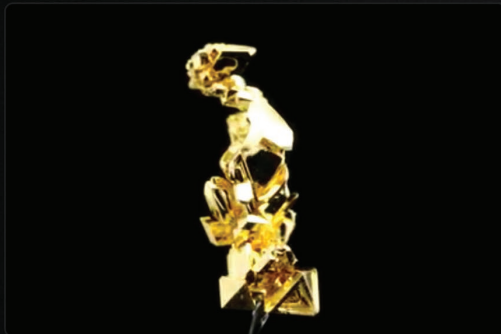
### 01. Overview

|               |                        |
|---------------|------------------------|
| Symbol        | Au                     |
| Atomic number | 79                     |
| Atomic weight | 197                    |
| Density       | 19.3 g/cm <sup>3</sup> |
| Melting point | 1064.18 °C             |
| Boiling point | 2856 °C                |
| Electrons     | 79                     |
| Protons       | 79                     |
| Neutrons      | 118                    |



### 02. Thermal properties

|                        |            |
|------------------------|------------|
| Phase                  | Solid      |
| Melting point          | 1064.18 °C |
| Boiling point          | 2856 °C    |
| Absolute melting point | 1337.33 K  |
| Absolute boiling point | 3129 K     |
| Critical pressure      | N/A        |
| Critical temperature   | N/A        |



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## Experiments



 You have found 0 of 11 secrets.

Chernobyl



The "Barking dog"



Cesium and water



Nitrogen triiodide



The Cannon



Rubber egg



The volcan



Fire snake

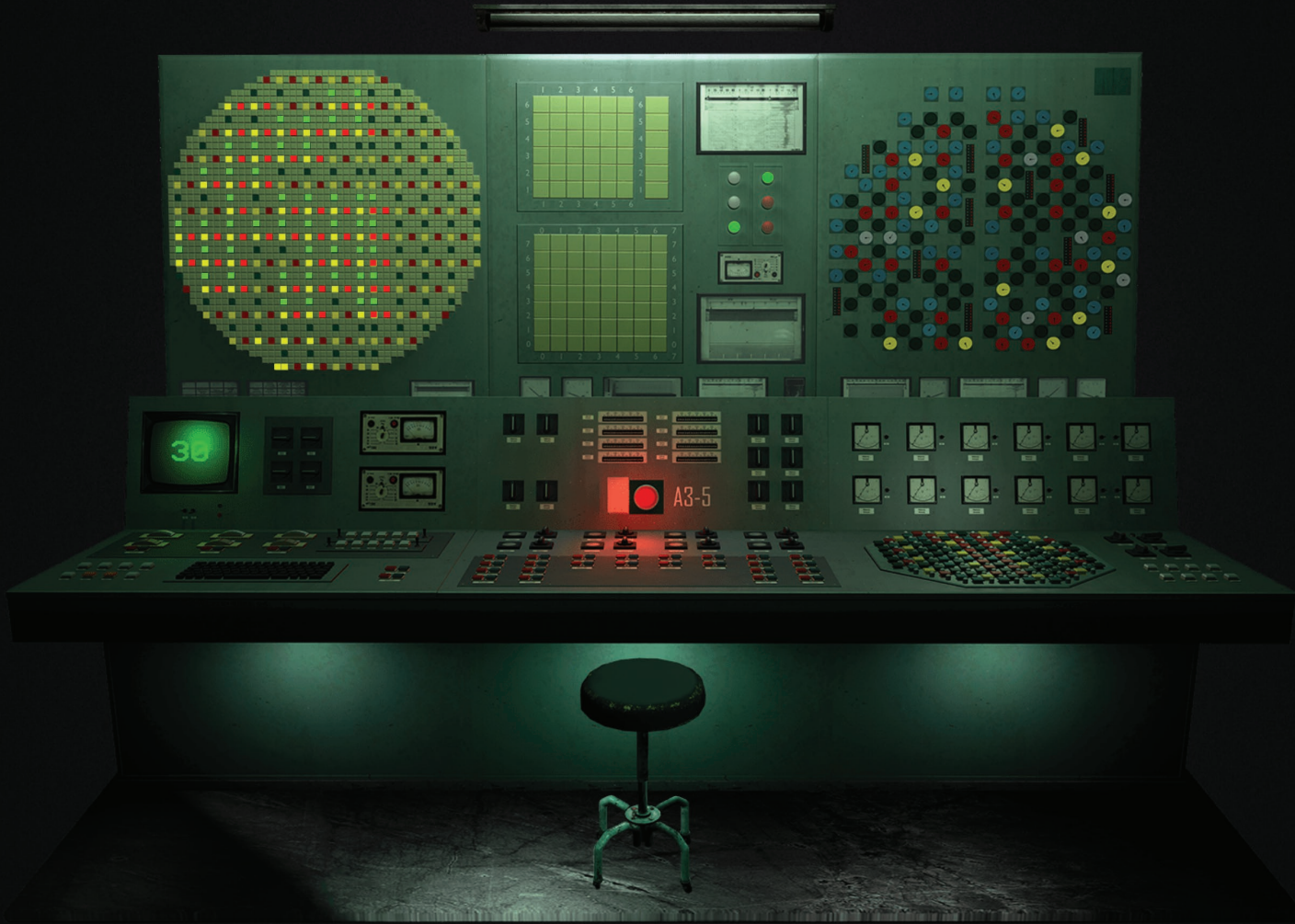


EN



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Chernobyl



EN





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The Cannon

Soft drink



Lollies



Soft drink + lollies

Help



In this experiment you can view the reaction between a soft drink and a bunch of lollies. To make the experiment, drag the elements that are pulsating into the glass at the center of the screen.

EN



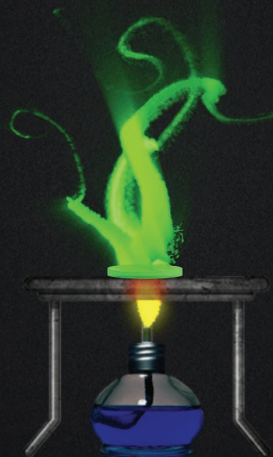
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Calcium gluconate and fire

Calcium gluconate



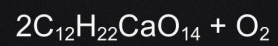
Fire



Help



In this experiment you can view the reaction between calcium gluconate and fire. To make the experiment, drag the elements that are pulsating into the glass at the center of the screen.



EN



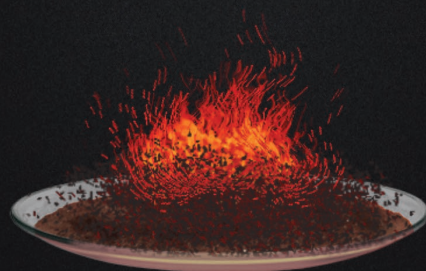


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The volcan

$\text{Cr}_2\text{H}_8\text{N}_2\text{O}_7$

Fire



$\text{Cr}_2\text{H}_8\text{N}_2\text{O}_7 + \text{O}_2$

Help



In this experiment you can view the reaction between ammonium dichromate and fire. To make the experiment, drag the elements that are pulsating into the glass at the center of the screen.

EN



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Hydrochloric acid and magnesium

Hydrochloric acid



Magnesium



$\text{HCl} + \text{Mg}$

Help



In this experiment you can view the reaction between the hydrochloric acid and magnesium solution. To make the experiment, drag the elements that are pulsating into the glass at the center of the screen.

EN





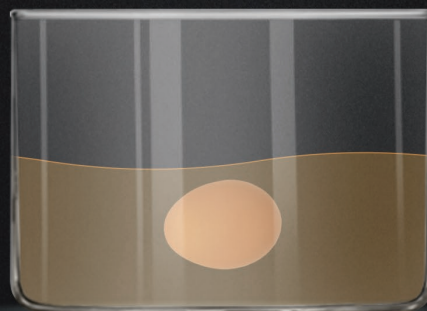
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Rubber egg

Vinegar



Boiled egg



$\text{CH}_3\text{COOH}$  + Boiled egg

Help



In this experiment you can view the reaction between the vinegar and one boiled egg. To make the experiment, drag the elements that are pulsating into the glass at the center of the screen.

EN



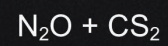
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The "Barking dog"

Nitrous oxide



Carbon disulfide



Help



In this experiment you can view the reaction between nitrogen monoxide and carbon disulfide. To make the experiment, drag the elements that are pulsating into the glass at the center of the screen.

EN







Periodic Table PRO is a reference application for each student, teacher, chemistry passionate and provides:

- A virtual laboratory with exciting packs of chemistry experiments
- Concise information for all chemical elements
- Multi-language support for English, Romanian, German, Russian, French, Italian, Spanish, Turkish, Chinese & Japanese
- 'In memoriam Chernobyl'
- Hidden app secrets
- Quick search & in-place information filtering
- Downloadable PDFs for each chemical element
- Presentation galleries & movie-clips

The project was awarded by Best Mobile App Awards, Awwwards & WebStock Awards:

- Platinum Award for Best Mobile Design
- Web Innovation Award
- Best Mobile App



**AWWARDS**

